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$C_2H_5OH(l) + 3 O_2(g) \rightarrow 2 CO_2(g) + 3 H_2O(l)$ $\Delta H_{rxn} = -1366.8 \text{ kJ/mol}$ A 100.0 mL Sample Of 0.300 M NaOH Is Mixed With A 100.0 mL Sample Of 0.305 M HNO_3 In A Coffee Cup Calorimeter. If Both Solutions Were Initially At 35.0 °C And The Temperature Of The Resulting Solution Was Recorded As 37.0 °C, Determine The ΔH_{rxn} (in Units Of KJ/mol). Apr 18th, 2024

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