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10 D. $2 / 8$ (or $1 / 4$) E. $2 / 10$ (or $1 / 5$) 2. The $\text{KClO}_3 / \text{O}_2$ Molar Ratio Is $2/3$. $2 \text{ Mol } \text{KClO}_3 / 3 \text{ Mol. } \text{O}_2 = 12.00 \text{ Mol } \text{KClO}_3 / X = 18.00 \text{ Mol. } X = 18.00 \text{ Mol Of } \text{O}_2$

3. Given The Following Equation: $2 \text{ K} + \text{Cl}_2 \rightarrow 2 \text{ KCl}$
How Many Grams Of KCl Is Produced From 2.50 G Of K And ... Mar 3th, 2024

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