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Take Home Exam, Ch 1 - 4 And Reaction Products Therefore The Ratio Is 1:1, So The Empirical Formula = HF B.  $(4.267\text{g UF}_6 \times 1 \text{ Mol UF}_6 \times 6 \text{ Mol F}) / (352.03\text{g} \times 1\text{mol}) = 0.07273 \text{ Mol F}$  Original Compound Gas:  $(0.0485 \text{ Mol F}) / (0.07273 \text{ Mol F}) = 66.7\%$  In Gas Or 922/1382 Solid: 33.3% In Solid Or 460/1382 C.  $(4.267\text{g UF}_6 \times 1 \text{ Mol UF}_6 \times 1$

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