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Because Each Formula Unit Of  $K_3PO_4$  Breaks Up On Dissolving Into Three  $K^+$  Ions  
And One  $PO_4^{3-}$  Ion:  $K_3PO_4(s) \div 3 K^+(aq) + PO_4^{3-}(aq)$  In Other Words, The  
Molality Of  $K^+$  Ions Is (3 Apr 8th, 2024)

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