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5.6 5.8 Kinetic Molecular Theory (KMT) Gas Particles Can Be Considered Point Particles Colliding With The Walls Of An Enclosure Randomly. The Higher The Temperature, The Faster Molecules Move, The More Kinetic Energy They Have. Kinetic Theory Gas Model For KM 4th, 2024

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13.1 Kinetic Theory And A Model For Gases The Word Kinetic

13.1 Kinetic Theory And A Model For Gases The Word Kinetic Refers To Motion. The Energy An Object Has Because Of Its

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CHAPTER 12 GASES AND KINETIC-MOLECULAR THEORY

3 Boyle's Law: The Volume-Pressure Relationship $\{V \propto 1/P$ Or $\{V = K (1/P)$ Or $PV = K\{P_1V_1 = K_1$ For One Sample Of A Gas. $\{P_2V_2 = K_2$ For A Second Sample Of A Gas. $\{k_1 = K_2$ For The Same Sample Of A Gas At The Same T. $\{Thus We Can Write Boyle's Law Mathematically As $P_1V_1 = P_2V_2$ 6th, 2024$

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