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2 $\text{KNO}_3 \rightarrow 2 \text{KNO}_2 + \text{O}_2$ C. 2 $\text{O}_3 \rightarrow 3 \text{O}_2$ D. $\text{NH}_4\text{NO}_3 \rightarrow \text{N}_2\text{O} + 2 \text{H}_2\text{O}$ E. $4 \text{CH}_3\text{NH}_2 + 9 \text{O}_2 \rightarrow 4 \text{CO}_2 + 10 \text{H}_2\text{O} + 2 \text{N}_2$ F. $\text{Cr}(\text{OH})_3 + 3 \text{HClO}_4 \rightarrow \text{Cr}(\text{ClO}_4)_3 + 3 \text{H}_2\text{O}$ W 1th, 2024.

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 $\text{Cr}_2(\text{SO}_4)_3 + 4 \text{ K}_2\text{SO}_4 + 3 \text{ I}_2 + 7 \text{ H}_2\text{O}$ A) How

Many Moles Of Potassium Dichromate ($K_2Cr_2O_7$) Are Required ... = 407.9 g AgBr This Is The Theoretical Yield Yields 22 b) Calculate The Percent Yield If 375.0 g Of Silver Bromide Was Obtained From The Reaction Theoretical Yield = 407.9 g AgBr Percent Yield = $100 \times \frac{\text{Actual Yield}}{\text{Theoretical Yield}}$ 2th, 2024.

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