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 Following Balanced Equation:  $3 \text{Ag(s)} + 4 \text{HNO}_3 \text{(aq)} \rightarrow 3 \text{AgNO}_3 \text{(aq)} + 2 \text{H}_2 \text{O(l)} + \text{NO} \dots$  15th, 2024.  
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Actual yield = 1.45g Theoretical yield = 5.95g  $\text{KClO}_4 \times 1 \text{ mol} = 122.55 \text{ g}$   
 $\times 2 \text{ O}_2 = 32.00 \text{ g} \times 1 \text{ mol} = 3.11 \text{ g O}_2$   
%Yield =  $\frac{1.45 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$  34.  $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$

Purity = 62.8%? pure  $\text{FeCO}_3$  8th, 2024 Practice Work 53 Stoichiometry-04 Mixed Stoichiometry

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