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A) [Co(H<sub>2</sub>O)<sub>6</sub>]<sup>2+</sup> Co + 2 E) [FeO]<sub>4</sub><sup>2-</sup> Fe + 6 I) K<sub>2</sub>[CoCl<sub>4</sub>] Co + 2 B) [CrCl<sub>6</sub>]<sup>3-</sup> Cr

+3 F)  $[\text{Mn}(\text{CN})_6]^{4-}$  Mn +2 J)  $\text{K}_3[\text{AuF}_6]$  Au +3 C)  $[\text{Co}(\text{NH}_3)_6]\text{Cl}_2$  Co +2 G)  $[\text{Ni}(\text{CO})_4]$  Ni 0 K)  $(\text{NH}_4)_2[\text{IrCl}_6]$  Ir +4 D)  $[\text{Co}(\text{NH}_3)_5\text{Cl}]\text{Cl}_2$  Co +3 H)  $[\text{Ni}(\text{EDTA})]^{2-}$  Ni +2 L)  $\text{Na}[\text{Mn}(\text{CO})_5]$  Mn -1 TASK 2 - Writing Half Equations A)  $2\text{I}^- \rightarrow \text{I}_2 + 2\text{e}^-$  - B ... Feb 29th, 2024

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Metal Carbonate  $\rightarrow$  Metal Oxide + Carbon Dioxide (on Heating)  $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$  TASK 4 - WRITING BALANCED EQUATIONS 1) Balance The Following Equations. Mar 30th, 2024

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(A2 BBoF 23) 4 Kc For The Equilibrium Below Is 7.50 At 50 °C. If 50.0 G Of Ethanol Is Mixed With 50.0 G Of Propanoic Acid, What Mass Of Ethyl Propanoate Will Be Formed At Equilibrium?  $\text{CH}_3\text{CH}_2\text{OH(l)} + \text{CH}_3\text{CH}_2\text{COOH(l)} \rightleftharpoons \text{CH}_3\text{CH}_2\text{COOCH}_2\text{CH}_3\text{(l)} + \text{H}_2\text{O(l)}$  (((((( Feb 9th, 2024

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Volume Used / Cm<sup>3</sup> 23.5 23.4 22.7 23.3 A Name The Piece Of Apparatus That Should Be Used To Measure The 25.0 Cm<sup>3</sup> Samples Of The Vinegar Into The Conical Flask. Pipette B Name The Piece Of Apparatus That Is Used For The Sodium Hydroxide Solution. Burette Jan 9th, 2024

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