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$11.6 \text{ Mol Of Ethanol In } (893.7 - 536.2) = 357.5 \text{ G Of Water. The Molarity Of Ethanol In The Solution is therefore } (11.6 \text{ mol}) \div (0.3575 \text{ kg}) = 32.4 \text{ mol kg}^{-1}.$ 4
Colligative properties: Escaping tendency of the solvent

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Page 4 Chem1 Chemical Bonding Stability And Reactivity The Fall In Energy When Atoms Join Together Is A Measure Of The Stability Of The New Aggregate. In Order To Be Regarded As A Molecule, The Aggregate Must Be Sufficiently Stable To Resist Disruption By Thermal Motions Long Enough To Enable The Observation Of Whate Jan 5th, 2023

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Circles Question Answer Mark Comments 7(a) $20(5)^3$

$+ k(5)^2 + 9(5) - = 0$ M1 $125 + +25k = 45 - 20$ 0 A1 k

$= -6$ A1 7(b) x^3 M1 $-6x^2 + 9x - 20 - (x \times 5)(2 + kx +$

4) Sight Of x^2 And 4 In A Quadratic Factor $(x - -5)(2x$

$+ 4)$ A17c Tests 'b² - 4ac' for The Quadratic M1 Ft

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