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(ii) Write A Chemical Equation For The Reaction Between Chlorine And Aqueous Potassium Bromide [1] (b) A Sample Of Vanadium Chloride Was Weighed And Dissolved In Water. An Excess Of Aqueous Sil 6th, 2024

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(d) The Oxidation States Of Vanadium In Its Compounds Are V(+5), V(+4), V(+3) And V(+2). The Vanadium(III) Ion Can Behave As A Reductant Or An Oxidant. (i) Indicate On The Following Equation Which Reactant Is The Oxidant. $2V^{3+} + Zn \rightarrow 2V^{2+} + Zn^{2+}$ [1] (ii) Which Change In The Following 2th, 2024

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