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Frequency Range Table Feb 9th, 2024Electrochemical  
Methods Solution ManualElectrochemical Methods For  
Neuroscience Provides An Updated Summary Of The  
Current, Albeit Evolving, State Of The Art And Lays The  
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 Methods Chapter 2 ... $N\text{O} + \text{ne} \rightarrow \text{NRR}$  (2.1.36) The Free  
 Energy Of The Cell Reaction Is Given By,  $\Delta G = \Delta G_0 +$   
 $RT \ln \frac{a_{\text{R}} a_{\text{N}}}{a_{\text{H}^+} a_{\text{O}}}$  (2.1.38)  
 Which Since  $\Delta G = -nFE_{\text{rxn}}$  And  $a_{\text{H}^+} = 1$   
 Becomes,  $E = E_0 + \frac{RT}{nF} \ln \frac{a_{\text{R}} a_{\text{N}}}{a_{\text{O}}}$ . (2.1.40)  
 This Expression Is Called The Nernst Equation. This  
 Relation Can Mar 1th, 2024Evaluation Of The Metal  
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