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(2.1.36) The Free Energy Of The Cell Reaction Is Given By,  $\Delta G = \Delta G^0 + RT \ln \frac{A_{N_2} A_{H^+} A_{O_2}}{A_{NH_3} A_{H_2}}$  (2.1.38) Which Since  $\Delta G = -nFE_{rxn}$  And  $A_{H^+} = A_{H_2} = 1$  Becomes,  $E = E^0 + \frac{RT}{nF} \ln \frac{A_{N_2} A_{O_2}}{A_{NH_3}}$ . (2.1.40) This Expression Is Called The Nernst Equation. This Relation Can Feb 29th, 2024

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