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(2.1.36) The Free Energy Of The Cell Reaction Is Given By, $DG = DG0 + RT \ln \frac{A_n R}{A_n H^+ H A_n O O A NH 2 H2}$ (2.1.38) Which Since $DG = NFE_{rxn}$ And $A H^+ = A H 2 = 1$ Becomes, $E = E0 + \frac{RT}{nF} \ln \frac{A_n R}{A_n O O}$. (2.1.40) This Expression Is Called The Nernst Equation. This Relation Can Feb 1th, 2024
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