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Paper 1 Physical Chemistry Multiple Choice Questions  $\Delta H / \text{KJ Mol}^{-1}$   $-822.0$   $0$   $0$   $\Delta \text{MolS} / \text{J K}^{-1}$   $-1$   $90.0$   $131.0$   $27.0$  The Standard Entropy Value For Steam Is A  $+332 \text{ J K}^{-1} \text{ Mol}^{-1}$  B  $+189 \text{ J K}^{-1} \text{ Mol}^{-1}$  C  $+145 \text{ J K}^{-1} \text{ Mol}^{-1}$  D  $+85 \text{ J K}^{-1} \text{ Mol}^{-1}$  (Total 1 Mark) Q25.Which One Of The Equations Below Represents A Reaction That Is Feasible ... Apr 13th, 2024

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