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0.05 Mol ... Feb 4th, 2024

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A) $[\text{Co}(\text{H}_2\text{O})_6]^{2+} + \text{Co} + 2\text{e}^-$ E) $[\text{FeO}_4]^{2-} + \text{Fe} + 6\text{e}^-$ I) $\text{K}_2[\text{CoCl}_4] + \text{Co} + 2\text{e}^-$ B) $[\text{CrCl}_6]^{3-} + \text{Cr} + 3\text{e}^-$ F) $[\text{Mn}(\text{CN})_6]^{4-} + \text{Mn} + 2\text{e}^-$ J) $\text{K}_3[\text{AuF}_6] + \text{Au} + 3\text{e}^-$ C) $[\text{Co}(\text{NH}_3)_6]\text{Cl}_2 + \text{Co} + 2\text{e}^-$ G) $[\text{Ni}(\text{CO})_4] + \text{Ni} + 0\text{e}^-$ K) $(\text{NH}_4)_2[\text{IrCl}_6] + \text{Ir} + 4\text{e}^-$ D) $[\text{Co}(\text{NH}_3)_5\text{Cl}]\text{Cl}_2 + \text{Co} + 3\text{e}^-$ H) $[\text{Ni}(\text{EDTA})]^{2-} + \text{Ni} + 2\text{e}^-$ L) $\text{Na}[\text{Mn}(\text{CO})_5] + \text{Mn} + 1\text{e}^-$ TASK 2 – Writing Half Equations A) $2\text{I}^- \rightarrow \text{I}_2 + 2\text{e}^-$ B ... Apr 1th, 2024

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Metal Carbonate → Metal Oxide + Carbon Dioxide (on Heating) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
TASK 4 – WRITING BALANCED EQUATIONS 1) Balance The Following Equations.
Apr 8th, 2024

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- Choose One Of The Testing Markers And Draw A Thick Line Near The Bottom Of The Filter Paper - About $\frac{1}{4}$ Inch From The Bottom.
- Pour A Small Amount Of Water Into The Large Cup And Then Hang The Paper Strip In The Apr 5th, 2024

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