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A) $[\text{Co}(\text{H}_2\text{O})_6]^{2+} + \text{Co}^{+2}$ E) $[\text{FeO}_4]^{2-} + \text{Fe}^{+6}$ I) $\text{K}_2[\text{CoCl}_4] + \text{Co}^{+2}$ B) $[\text{CrCl}_6]^{3-} + \text{Cr}^{+3}$ F) $[\text{Mn}(\text{CN})_6]^{4-} + \text{Mn}^{+2}$ J) $\text{K}_3[\text{AuF}_6] + \text{Au}^{+3}$ C) $[\text{Co}(\text{NH}_3)_6]\text{Cl}_2 + \text{Co}^{+2}$ G) $[\text{Ni}(\text{CO})_4] + \text{Ni}^0$ K) $(\text{NH}_4)_2[\text{IrCl}_6] + \text{Ir}^{+4}$ D) $[\text{Co}(\text{NH}_3)_5\text{Cl}]\text{Cl}_2 + \text{Co}^{+3}$ H) $[\text{Ni}(\text{EDTA})]^{2-} + \text{Ni}^{+2}$ L) $\text{Na}[\text{Mn}(\text{CO})_5] + \text{Mn}^{-1}$ TASK 2 – Writing Half Equations A) $2\text{I}^- \rightarrow \text{I}_2 + 2\text{e}^-$ B ...
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Metal Carbonate \rightarrow Metal Oxide + Carbon Dioxide (on Heating) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
TASK 4 – WRITING BALANCED EQUATIONS 1) Balance The Following Equations.
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