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Mixed Stoichiometry Problems 1. How Many Moles Of  $\text{H}_2$  Would Be Required To Completely React With  $\text{O}_2$  To Produce 5 Moles Of Water? 5 Mol  $\text{H}_2$  2.  $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$  A. Balance This Equation B. What Mass Of  $\text{H}_2\text{SO}_4$  Would Be Required To React With 0.75 Mol Of  $\text{NaOH}$ ? 37g 3. What Mass Of  $\text{NO}_2$  Is Formed When  $\text{NO}$  Reacts With 384 G Of  $\text{O}_2$ ? Apr 2th, 2024

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