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Required? $15.0g N_2O_4 \times \frac{1 \text{ Mol } N_2O_4}{92.0g N_2O_4} = 0.163 \text{ Mol } N_2O_4$
 $0.163 \text{ Mol } N_2O_4 \times \frac{2 \text{ Mol } O_2}{1 \text{ Mol } N_2O_4} = 0.326 \text{ Mol } O_2$ B. If 4.0×10^{-3} Moles Of Oxygen Reacted, How Many Grams Of N_2
Were Needed? $4.0 \times 10^{-3} \text{ Mol } O_2 \times \frac{1 \text{ Mol } N_2}{2 \text{ Mol } O_2} = 2.0 \times 10^{-3} \text{ Mol } N_2$... Mar 12th, 2024.

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Factor, We Determine The Mass That This Number Of Moles Of SO₃ Has. The First
Step Resembles The Exercises We Did In Section 6.4 "Mole-Mole Relationships In
Chemical Reactions". As Usual, We Start With The Quantity We Were Given: 3.59
 $\text{Mol } Fe_2O_3 \times \frac{3 \text{ Mol } SO_3}{1 \text{ Mol } Fe_2O_3} = 10.77 \text{ Mol } SO_3$ The Mol Fe_2O_3 Units Cancel,
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