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Assigning Oxidation Numbers Directions: Assign Oxidation Numbers To All Of The Elements In Each Of The Compounds Or Ions Below. 1. HCl 11. H 2SO 3 2. KNO 3 12. H 2SO 4 3. OH - 13. BaO 2 4. Mg3N2 14. KMnO 4 5. KClO 3 15. LiH 6. Al(NO 3)3 16. MnO 2 7. S 8 17. OF 2 8. H 2O2 18. SO 3 9. PbO 2 19. NH 3 10. NaHSO 4 20. Na Chemistry IF8766 Page 91 ... May 1th, 2024

#### AP CHEMISTRY NOTES 2-1 ASSIGNING OXIDATION ...

ASSIGNING OXIDATION NUMBERS RULES FOR ASSIGNING OXIDATION NUMBERS: 1. The Oxidation Number Of Any Free Element (including Diatomic Elements) Is Always 0. 2. The Sum Of The Oxidation Number Of All Of The Atoms In A Compound Is Always Equal To 0. 3. The Sum Of The Oxidation Numbers Of All Of Th Apr 4th, 2024

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4. The Sum Of The Oxidation Numbers Of All The Elements In A Polyatomic Ion Equals The Charge On The Ion. Directions: In The Following Questions, Give The Oxidation Number Of The Indicated Atoms/ion. 1. N In N 2 O 3 N3+ 2. S In H 2 SO 4 S6+ 3. C 2C0 4. C In CO C2+ 2+ 5. Na In NaCl Na+ 6. H I Mar 4th, 2024

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2. For Monoatomic Ions, The Oxidation Number Is The Same As The Charge. 3. All Metals In Group 1 Form +1 Ions And Group

2 Form +2 Ions When Found In A Compound. 4. When Found In A Compound Fluorine Has A - 1 Oxidation State. 5. Hydrogen Has A +1 Oxidation State When Bonded To A Non Jul 4th, 2024

#### **RULES FOR ASSIGNING OXIDATION NUMBERS**

6. The Algebraic Sum Of The Oxidation Numbers Of Elements In A Compound Is Zero. CaF 22 H SO 4+2+2(-1)=0 2(+1) + S + 4(-2) = 0 Find S = +6 For Ox. No. Of S 7. The Algebraic Sum Of The Oxidation Numbers Apr 4th, 2024

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