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14. Writing The Formulas Of Compounds. A. Binary And Ternary Ionic Compounds : Basically You Have To Work Backward From The Name To The Symbols. For Example Sodium Chloride Involves Sodium And Chlorine And The Formula Is NaCl. What You Need To Be Careful Of Is The Ratio Of Ions In The Compound. The Compound Must Be Neutral (have Equal Positive And Negative Charges). Feb 10th, 2024

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2. Write The Formulas For The Compounds Formed From These Pairs Of Ions. C. Ca^{2+} and K^{+} E. O^{2-} and Cu^{+} H. Sn^{4+} and Cl^{-} 3. Name The Following Binary Ionic Compounds. A. MnO_2 B. CaCl_2 D. SrBr_2 E. FeCl_3 Chapter Jan 1th, 2024

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9.5 The Laws Governing Formulas And Names • The Ways That Compounds Form Are Summed Up In Two Laws: The Law Of Definite Proportions And The Law Of Multiple Proportions. • To Name A Compound Or Write Its Formula, Follow The Flowcharts In Figures 9.20 And 9.22 To The Correct Name O Jan 4th, 2024

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