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$\text{Rate}_1 = k[\text{D}]^1 [\text{X}]^1 [\text{Y}]$ $\text{Rate}_2 = k[\text{D}]^2 [\text{X}]^2 [\text{Y}]$ Equation 5 In This Equation k Cancels Out. $[\text{D}]^1 = [\text{D}]^2$ Because We Are Starting Off With The Same Initial Concentration Of A In Each Trial. The Order Of Species D Which Is X Is Also Constant. Canceling Terms We Have Left: $\text{Rate}_1 = k[\text{B}]^1 [\text{Y}]$ $\text{Rate}_2 = k[\text{B}]^2 [\text{Y}]$ Equation 6 Consider The Following Chemical Reaction: 9th, 2024

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Very Rapidly. Also, The More Bonds Between Reacting Atoms In A Molecule, The
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Into Its “bits”. Substances With Strong Bonds (larger Bond Energies) Will React
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So, To Account For Effective Collisions, Another Factor, P Called The Probability Or
Steric Factor Is Introduced. So, $\text{Rate} = PZ \frac{AB}{RT} e^{-E_a/RT}$ The Activated Complex
Theory Or Transition State Theory Reactants \rightleftharpoons Activated Complex \rightarrow Products This
Theory Is Based On The Fact That 8th, 2024

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2 and Cl_2 to form HCl . 7 Chemistry-I. Figure 1.1: Plot Of Concentration Against T For A Reaction $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$ The Rate Of This Reaction Does Not Depend On Concentration. 1.6 First Order Kinetics A Reaction Of The first Order Is Represented As $X \rightarrow Y$ Where X Is The Reactant And Y The Product. The Rate Of The Reaction Will Be Directly 6th, 2024

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Texts: Silbey, Alberty, And Bawendi, Physical Chemistry, 4th Edition. Solutions Manual For Physical Chemistry, 4th Edition. Prerequisites: 'C' Or Better In CH 117; MA 126; PH 201 Or PH 221 . Strongly Recommend: CH 355 (Quantitative Analysis) And PH 221 . Course Objectives: ... 15th, 2024

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B. Transfer Translational KE To Vibrational Energy To Weaken/break Reactant Bonds
Molecules At A Given Temperature Possess A KE Distribution: Number Of Molecules
KE $E \geq E_A$, Activation Energy - Energy Barrier That Molecules Have To Surmount In
Order To React. Energy Is Needed To Break Reactant Bonds (endothermic Process).
14th, 2024

Chapter 14 Chemical Kinetics

3 Chemical Kinetics Factors That Affect Reaction Rates • Physical State Of The
Reactants In Order To React, Molecules Must Come In Contact With Each Other. If
The Reaction Is Happening Between A Solid And A Liquid It Will React Only On The
Surface. The More Homogeneous The Mixture Of Reactants, The Faster The
Molecules Can React. 13th, 2024

Chemical Kinetics Reaction Rates

5 The Overall Order Of A Reaction Is The Sum Of The Individual Orders: Rate ($M s^{-1}$)
 $= k[A][B]^{1/2}[C]^2$ Overall Order: $1 + \frac{1}{2} + 2 = 3.5 = \frac{7}{2}$ Or Seven-halves Order
Note: When The Order Of A Reaction Is 1 (first Order) No Exponent Is Written. Units
For The Rate Constant: The Units Of A Rate Constant Will Change Depending Upon

The Overall 18th, 2024

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