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Stoichiometry Review Worksheet Name: ___KEY___ 33. $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$

Actual yield = 1.45g Theoretical yield = 5.95g $\text{KClO}_4 \times 1 \text{ mol } 122.55 \text{ g} \times 2 \text{ O}_2 \times 1 \text{ KClO}_4 \times$

$32.00 \text{ g } 1 \text{ mol } = 3.11 \text{ g O}_2$ %Yield = $\frac{1.45 \text{ g } 3.11 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$ 34. $4\text{FeCO}_3 + \text{O}_2$

$\rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$ Percent purity = 62.8%? g pure FeCO_3 Feb 2th, 2024 Honors

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Calcium $\text{C}_2\text{O}_2 + 3 \text{Cr}(3)$ When The Potassium Chlorate (KClO_3) Is Heated, It Decomposes To Form Gases KCl And Oxygen (O_2).
 $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$ C. Feb 9th, 2024
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Stoichiometry Worksheet #1 Answers - My Chemistry Class
Stoichiometry Worksheet #1 Answers
1. Given The Following Equation: $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$, Show What The Following Molar Ratios Should Be.
A. $\text{C}_4\text{H}_{10} / \text{O}_2$ B. O_2 / CO_2 C. $\text{O}_2 / \text{H}_2\text{O}$ D. $\text{C}_4\text{H}_{10} / \text{CO}_2$ E. $\text{C}_4\text{H}_{10} / \text{H}_2\text{O}$
2. Given The Following Equation: $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$ A. How Many Moles Of O_2 Can Be Produced By ...
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