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Thereby Making The Reaction Irreversible. The Concentration-time D Jan 19th, 2024  
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Temperature We Will Once Again Look At The Formation Of Ammonia (NH<sub>3</sub>) From  
Nitrogen And Hydrogen (see Section Chemical Equilibrium Of The Thermodynamics  
Chapter). This Reaction Follows The Equation:  $N_2 + 3H_2 \rightleftharpoons 2NH_3$  (1)  $H_0 = 92 \text{ KJ Mol}$   
 $S_0 = 192 \text{ J Mol K}$  Jan 13th, 2024.

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15% Exam 3 15% Final Exam 20% Research Paper 10% Homework 20% Class  
Participation 5% Last Year, The Following Grading Scale Was Used:  $A \geq 92$   $A/B \geq 89$   
 $B \geq 82$   $B/C \geq 80$   $C \geq 71$   $C/D \geq 69$   $D \geq 61$   $F \leq 60$  If Necessary, The Grading Scale  
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