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$= k[D]^1 [B]^1$   $Y$   $k[D]^2 [B]^2$   $Y$  Equation 5 In This Equation  $k$  Cancels Out.  $[D]^1 = [D]^2$  Because We Are Starting Off With The Same Initial Concentration Of  $A$  In Each Trial. The Order Of Species  $D$  Which Is  $X$  Is Also Constant. Canceling Terms We Have Left:  $Rate_1$   $Rate_2 = k[B]^1$   $Y$   $k[B]^2$   $Y$  Equation 6 Consider The Following Chemical Reaction: 2th, 2024

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Laws PLAN: We Inspect The Exponents In The Rate Law, Not The Coefficients Of The Balanced Equation, To Find The Individual Orders. We Add The Individual Orders To Get The Overall Reaction Order. (a) The Exponent Of [NO] Is 2 And The Exponent Of [O<sub>2</sub>] Is 1, So The Reaction Is Second 2th, 2024

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2and Cl<sub>2</sub> to Form HCl. 7 Chemistry-I. Figure 1.1: Plot Of Concentration Against T For A Reaction  $H_2 + Cl_2 \rightarrow 2HCl$  The Rate Of This Reaction Does Not Depend On Concentration. 1.6 First Order Kinetics A Reaction Of The first Order Is Represented As  $X \rightarrow Y$  Where X Is The Reactant And Y The Product. The Rate Of The Reaction Will Be Directly 1th, 2024

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B. Transfer Translational KE To Vibrational Energy To Weaken/break Reactant Bonds Molecules At A Given Temperature Possess A KE Distribution: Number Of Molecules KE  $E \geq E_A$ , Activation Energy - Energy Barrier That Molecules Have To Surmount In Order To React. Energy Is Needed To Break Reactant Bonds (endothermic Process). 6th, 2024

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