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 And Dependence On Temperature We Will Once Again  
 Look At The Formation Of Ammonia ( $NH_3$ ) From  
 Nitrogen And Hydrogen (see Section Chemical  
 Equilibrium Of The Thermodynamics Chapter). This  
 Reaction Follows The Equation:  $N_2 + 3H_2 \rightleftharpoons 2NH_3$  (1)

$H_0 = 92 \text{ kJ/mol}$   $S_0 = 192 \text{ J/mol K}$  1th, 2024  
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Exam 1 15% Exam 2 15% Exam 3 15% Final Exam  
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L.35 In The Textbook Question: In The Third Reaction Step The Compound "FeBr<sub>2</sub>" Needs To Be "Fe<sub>3</sub>Br<sub>8</sub>".

1A.15 Typo. Answer Is  $N = 3$  To  $N = 1$  As Energy Is Emitted. 1B.27 Solution Manual States  $\Delta v = 5.0 \text{ M.s}^{-1}$  Velocity Was Given As  $5.00 \pm 5.0 \text{ M.s}^{-1}$  1th,

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