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The $\text{KClO}_3 / \text{O}_2$ Molar Ratio Is $2/3$. 2 Mol $\text{KClO}_3 / 3$ Mol. $\text{O}_2 = 12.00$ Mol $\text{KClO}_3 / X = 18.00$ Mol. $X = 18.00$ Mol Of O_2 3. Given The Following Equation: $2 \text{K} + \text{Cl}_2 \rightarrow 2 \text{KCl}$ How Many Grams Of KCl Is Produced From 2.50 G Of K And ... Mar 11th, 2024.

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Worksheet Name: KEYChemistry 11 Stoichiometry Review Worksheet

Name: ___KEY___ 33. $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$ Actual yield = 1.45g Theoretical yield = 5.95g $\text{KClO}_4 \times 1\text{mol} \times 122.55\text{g} \times 2 \times 2 = 1\text{KClO}_4 \times 32.00\text{g} \times 1\text{mol} = 3.11\text{g}$ O_2 %Yield = $\frac{1.45\text{g}}{3.11\text{g}} \times 100\% = 46.6\%$ 34. $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$

Percent purity = 62.8%?g pure FeCO_3 May 17th, 2024.

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