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The Number Of Moles Of Compound And The Molar Heat Of Combustion So By
Multiplying Them Together We Can Find The Heat Generated. $(0.875 \text{ Mol})(1350$
 $\text{KJ/mol}) = 1180 \text{ KJ} = 118000 \text{ J}$ Plug This Value Into Our Equation And Then Rearrange
To Find The Change In The Feb 18th, 2024

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