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T = 110.0 4.18 3.818103 G J Mc Q = 8.30 C T F = T I + T = 21.4 + 8.30 = 29.7 C 7) In Order To Write The Thermochemical Equation, You First Need To Calculate The MOLAR ENTHALPY CHANGE. I) Calculate The Heat Change, Q, For The Surroundings – In This Case, The Surroundings Is The TOTAL VOLUME OF THE SOLUTION 4th, 2024

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The Increase In Temperature Be? We Are Given Both The Number Of Moles Of Compound And The Molar Heat Of Combustion So By Multiplying Them Together We Can Find The Heat Generated. (0.875 Mol)(1350 KJ/mol) = 1180 KJ = 118000 J Plug This Value Into Our Equation And Then Rearrange To Find The Change In The 2th, 2024

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Calorimetry Practice Worksheet (Section 17.2) 1.) Compound A Is Burned In A Bomb Calorimeter That Contains 2.50 Liters Of Water. If The Combustion Of 0.175 Moles Of This Compound Causes The Temperature Of The Water To Rise 45.00C, What Is The Molar Heat Of Combustion Of Compound A (in KJ/ 3th, 2024

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O) (4.184 J/g °C)(55.4 °C) H = 1310 KJ This Is The Amount Of Energy Generated When 0.315 Moles Of Hexane Is Burned. To

Find The Molar Heat Of Combustion: H = -1310 KJ/0.315 Moles = - 4160 KJ/mol. 2) If 22.0 Grams Of Propane (C 3 H 8) Is Combusted In A Bomb Calorimeter Containing 3.25 L 3th, 2024

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